

tion just short of that necessary to satisfy the chlorine present. Of course if the chlorine ran so high as to render concentration unnecessary, then 100 cc. of the water itself with the indicator and the partial dose of silver nitrate would be the proper contents for the comparison dish. By these means the eye is greatly aided in noting the slightest appearance of red tint, for in respect of turbidity both dishes are alike. The results are very satisfactory.—*W. P. Mason, Rensselaer Polytechnic Institute, January 20, 1894.*

*The Phenolsulphonic Acid Process for Nitrates in Water.*—The interference of chlorides with this process, resulting in readings decidedly lower than the truth, is well known, but the method of determination is so easy and convenient, that it occurred to me to try the addition of sodium chloride to the comparison standards rather than abandon the process.

The "chlorine" in the water under examination having been previously determined, an appropriate volume of standardized sodium chloride solution is added to each evaporation of standard potassium nitrate solution. Thus the water to be examined, and the nitrate solutions with which it is compared, all contain the same quantity of chlorine. The results are very satisfactory.—*W. P. Mason, Rensselaer Polytechnic Institute, Jan. 18, 1894.*